Properties of Matter

Substance VS Mixture

1. \_\_\_\_\_\_\_\_\_\_sodium
2. \_\_\_\_\_\_\_\_\_\_soil
3. \_\_\_\_\_\_\_\_\_\_coffee
4. \_\_\_\_\_\_\_\_\_\_oxygen
5. \_\_\_\_\_\_\_\_\_\_70% isopropyl alcohol
6. \_\_\_\_\_\_\_\_\_\_carbon dioxide
7. \_\_\_\_\_\_\_\_\_\_cake batter
8. \_\_\_\_\_\_\_\_\_\_chicken noodle soup
9. \_\_\_\_\_\_\_\_\_\_iron
10. \_\_\_\_\_\_\_\_\_\_salt water
11. \_\_\_\_\_\_\_\_\_\_chocolate chip ice cream
12. \_\_\_\_\_\_\_\_\_\_nitrogen
13. \_\_\_\_\_\_\_\_\_\_eggs
14. \_\_\_\_\_\_\_\_\_\_blood
15. \_\_\_\_\_\_\_\_\_\_table salt
16. \_\_\_\_\_\_\_\_\_\_nail polish
17. \_\_\_\_\_\_\_\_\_\_milk
18. \_\_\_\_\_\_\_\_\_\_soda

Homogeneous VS Heterogenous OR element VS compound

Using the same items from above label them as homogenous or heterogenous (if it is a mixture) OR element or compound (if it is a substance)

1. \_\_\_\_\_\_\_\_\_sodium
2. \_\_\_\_\_\_\_\_\_\_soil
3. \_\_\_\_\_\_\_\_\_\_coffee
4. \_\_\_\_\_\_\_\_\_\_oxygen
5. \_\_\_\_\_\_\_\_\_\_70% isopropyl alcohol
6. \_\_\_\_\_\_\_\_\_\_carbon dioxide
7. \_\_\_\_\_\_\_\_\_\_cake batter
8. \_\_\_\_\_\_\_\_\_\_chicken noodle soup
9. \_\_\_\_\_\_\_\_\_\_iron
10. \_\_\_\_\_\_\_\_\_\_salt water
11. \_\_\_\_\_\_\_\_\_\_chocolate chip ice cream
12. \_\_\_\_\_\_\_\_\_\_nitrogen
13. \_\_\_\_\_\_\_\_\_\_eggs
14. \_\_\_\_\_\_\_\_\_\_blood
15. \_\_\_\_\_\_\_\_\_\_table salt
16. \_\_\_\_\_\_\_\_\_\_nail polish
17. \_\_\_\_\_\_\_\_\_\_milk
18. \_\_\_\_\_\_\_\_\_\_soda

Physical change VS Chemical Change

1. \_\_\_\_\_\_\_\_\_\_\_\_\_sodium hydroxide dissolves in water
2. \_\_\_\_\_\_\_\_\_\_\_\_\_hydrochloric acid reacts with potassium hydroxide to produce a salt, water, and heat
3. \_\_\_\_\_\_\_\_\_\_\_\_\_a pellet of sodium is sliced in two
4. \_\_\_\_\_\_\_\_\_\_\_\_\_water is heated and changed to steam
5. \_\_\_\_\_\_\_\_\_\_\_\_\_potassium chlorate decomposes to potassium chloride and oxygen gas
6. \_\_\_\_\_\_\_\_\_\_\_\_\_iron rusts
7. \_\_\_\_\_\_\_\_\_\_\_\_\_when placed in water, a sodium pellet catches on fire as hydrogen gas is liberated and sodium hydroxide forms
8. \_\_\_\_\_\_\_\_\_\_\_\_\_evaporation
9. \_\_\_\_\_\_\_\_\_\_\_\_\_ice melting
10. \_\_\_\_\_\_\_\_\_\_\_\_\_milk sours
11. \_\_\_\_\_\_\_\_\_\_\_\_\_sugar dissolves in water
12. \_\_\_\_\_\_\_\_\_\_\_\_\_wood rotting
13. \_\_\_\_\_\_\_\_\_\_\_\_\_egg frying
14. \_\_\_\_\_\_\_\_\_\_\_\_\_tree growing
15. \_\_\_\_\_\_\_\_\_\_\_\_\_food digested
16. \_\_\_\_\_\_\_\_\_\_\_\_\_koolaid in water
17. \_\_\_\_\_\_\_\_\_\_\_\_\_ gasoline burns in a car engine
18. \_\_\_\_\_\_\_\_\_\_\_\_\_sublimation of moth balls
19. \_\_\_\_\_\_\_\_\_\_\_\_\_liquefying oxygen
20. \_\_\_\_\_\_\_\_\_\_\_\_\_melting of ice cream
21. \_\_\_\_\_\_\_\_\_\_\_\_\_formation of dew on grass

Physical properites VS Chemical properties

If it is physical put a P AND extensive or intensive; If it is chemical put a C.

58\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_sulfur is a bright yellow solid

59. \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_sulfur has a low melting point

60. \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_sulfur causes silver to tarnish

61. \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_aluminum is very malleable

62. \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_monuments made of copper corrode in acid rain

63. \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_copper is a good conductor of electricity

64. \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_color

65. \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_reactivity

66. \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_flammability

67. \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_odor

68. \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_solubility

69.\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_density

70. \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_conductivity