

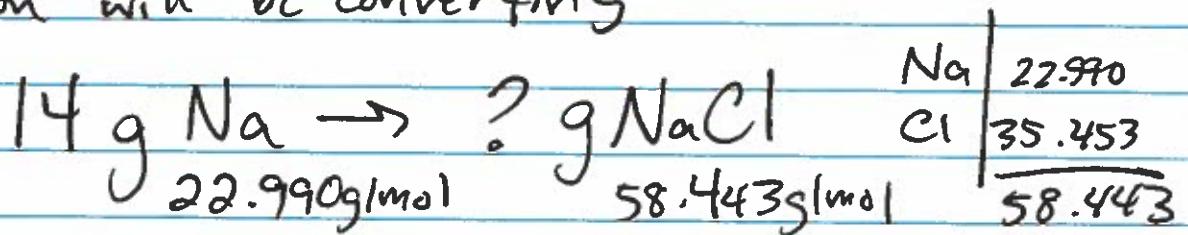
Mixed Stoich Notes

- Now use mol ratios to convert from g of sub A to g sub B

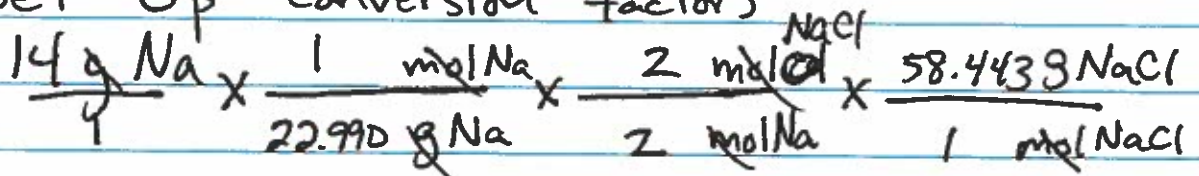


1st * Make sure the eq is balanced

2nd * Determine the molar mass of the subs you will be converting

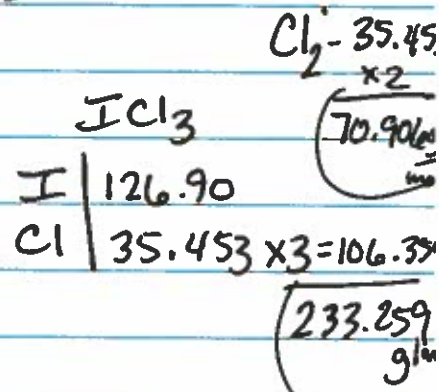
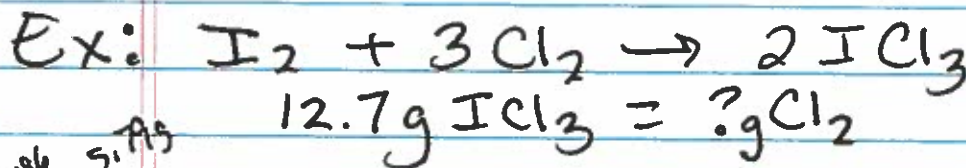


3rd * Set up conversion factors



4th * fill in conv. factors & solve

$$\frac{1636.404}{45.98} = \boxed{35.59 \text{ g NaCl}}$$



$$\frac{12.7 \text{ g ICl}_3}{1} \times \frac{1 \text{ mol ICl}_3}{233.259 \text{ g ICl}_3} \times \frac{3 \text{ mol Cl}_2}{2 \text{ mol ICl}_3} \times \frac{70.906 \text{ g Cl}_2}{1 \text{ mol Cl}_2} = \boxed{5.79 \text{ g Cl}_2}$$

$$\frac{2,701.5186}{466.518} = 5.79 \text{ g}$$