HOW TO COUNT ATOMS

Background Information:

- <u>subscripts</u> little numbers that tell how many atoms there are (ex: In 3H₂O, the ₂ is the subscript)
- <u>coefficients</u> regular-sized numbers that tell how many molecules there are (ex: In 3H₂O, the **3** is the coefficient)

Example: 3H₂O

- The subscript 2 in the example above comes after the H.
 This means there are two H's (hydrogen atoms) in each molecule.
- The coefficient 3 shows that there are three of the H₂O molecules.

When counting atoms in the example $(3H_2O)$, we can look at the atoms this way [since there are 3 of the H₂O molecules]: (3 water molecules each with 2 hydrogen and 1 oxygen)



So now, we can count the number of H's and the number of O's.
 Thus, 3H₂O has 6 hydrogen atoms and 3 oxygen atoms.

However, drawing out each equation is not practical, so...

Although the 1 is usually not written, $3H_2O$ can be written as $3H_2O_1$. (In other words, $3H_2O$ and $3H_2O_1$ are the same thing.) The number of atoms can be counted without drawing everything out <u>by multiplying the</u> <u>coefficient by all the subscripts in the molecule (you get the number of</u> <u>atoms)</u>. So with the case of $3H_2O_1$, we can multiply the 3 with the ₂ to find that there are 6 H's. Then we can multiply the 3 with the ₁ to find that there are 3 O's.

HOW TO COUNT ATOMS IN A CHEMICAL FORMULA

(5 Easy Steps)

Step 1: Write the chemical formula

Step 2: List all the atoms

Step 3: Count the number of atoms of each element in ONE molecule.

Step 4: Multiply the number of atoms of each by the coefficient.

Step 5: Make sure you answer makes sense.

PRACTICE TIME: Count the number of each type of atom in each molecule.

P ₄ + O ₂		4P ₂ O ₃
Ρ	4	P 2*4=8
0	2	O 3*4=12

Two More Practice Example

 $6 C_2 H_5 OH$ (Ethanol)

H - 6*6=<mark>36</mark> O - 1*6=<mark>6</mark>

4 C₆H₁₂O₆ (Glucose - Sugar) C - 6*4=<mark>24</mark> H - 12*4=<mark>48</mark> O - 6*4=<mark>24</mark>